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 Equations Practice
 Problems Chapter 16

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 1 6 **شاطر** Calorimetry
 Examples: How to Find
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Chapter 6 Chapter 6 -
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Formulas - Lecture Review

Practice Problems

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Energy is the essence of our very existence as individuals and as a

society. The food that we eat furnishes the energy to live, work, and play, just as the coal and oil consumed by manufacturing and transportation systems power our modern industrialized civilization. In the past, huge quantities of carbon-based fossil fuels have been available for the taking.

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Section 6.1 Notes

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Enthalpy and

Calorimetry Section 6.2

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6.3 Hess's Law Section

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6.4 Enthalpy of Formation Section 6.4
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Section 6.2 Enthalpy and Calorimetry
Enthalpy is a thermodynamic quantity that accounts for heat flow during the course of a chemical reaction Equals the energy contained by the system as well as the pressure/volume work done on or by the system: $H = E + PV$

Thermochemistry
CHAPTER 6:
THERMOCHEMISTRY
168 6.20 Strategy: The work done in gas expansion is equal to the product of the

external, opposing pressure and the change in volume. $w = -P\Delta V$ We assume that the volume of liquid water is zero compared to that of steam.

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Section 6.2. Enthalpy and Calorimetry. A 100.0 g sample of water at 90 °C is added to a 100.0 g sample of water at 10 °C. The final temperature of the water is: a) Between 50 °C and 90 °C b) 50 °C c) Between 10 °C and 50 °C.

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Thermochemistry -
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Thermochemistry.
Chapter 6. Energy is the capacity to do work.
Radiant energy .

comes from the sun and is earth's primary energy source. Thermal energy. is the energy associated with the random motion of atoms and molecules. Chemical energy. is the energy stored within the bonds of chemical substances.

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