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the motion of particles of a gas
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The three fundamental gas laws discover the relationship of pressure, temperature, volume and amount of gas. Boyle's Law tells us that the volume of gas increases as the pressure decreases. Charles' Law tells us that the volume of gas increases as the temperature increases. And Avogadro's Law tell us that the volume of gas increases as the amount of gas increases. The ideal gas law is the combination of the three simple gas laws.

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Gases - Section 5 of General Chemistry Notes is 18 pages in length (page 5-1 through page 5-18) and covers ALL you'll need to know on the following lecture/textbook topics:. SECTION 5 - Gases 5-1 -- Atmospheric Pressure · The Barometer · Two Factors Affecting Barometric Pressure 5-2 -- Units of Pressure · mm Hg, torr, atm, Pa, psi 5-2 - Boyle's Law

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General gas equation The ideal gas law is the equation of state of a hypothetical ideal gas. This equation was first stated by French engineer and physicist Emile Clapeyron (1799-1864) in 1834 as a combination of three empirical gas laws proposed by Robert Boyle, Joseph Louis Gay-Lussac, and Amedeo Avogadro.

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