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a A + b B f x Xwith a catalyst C. Rate = k [A]m[B]n[C]p. The exponents m, n, and p •are the reaction order •can be 0, 1, 2 or fractions •must be determined by experiment! 12. © 2009 Brooks/Cole - Cengage. Interpreting Rate Laws. Rate = k [A]m[B]n[C]p. •If m = 1, rxn. is 1st order in A Rate = k [A]1.

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Chemistry Chemistry: Principles and Reactions When tin comes in contact with the oxygen in the air, tin(IV) oxide, SnO 2 , is formed. Sn (s) +O 2 (g) → SnO 2 (s) A piece of tin foil, 8.25 cm × 21.5 cm × 0.600 mm (d = 7.28 g / cm 3) is exposed to oxygen.

When tin comes in contact with the oxygen in the air, tin ...

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