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5 The Periodic Law

Consider the neutral atom with 53 protons and 74 neutrons to answer the following questions. 53. a. What is its atomic number? 127. b. What is its mass number? atomic number. c.

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atomic mass = p+n. there are 18 columns in the periodic table; each has a group number. give the group numbers that make up each of the following blocks: a) s block. b) p block. c) d block. a) s block: 1-2. b) p block: 13-18. c) d block: 3-12. when an electron is added to a neutral atom, energy is.

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When an electron is added to a neutral atom, energy is . (a) always absorbed (c) either absorbed or released (b) always released (d) burned away 2.

Chapter 5 Review The Periodic Law Section 1

CHAPTER 5 REVIEW The Periodic Law SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. c In the modern periodic table, elements are ordered (a) according to decreasing atomic mass. (b) according to Mendeleev's original design. (c) according to increasing atomic number. (d) based on when they were discovered. 5 The Periodic Law

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Question: A Review Constants Periodic Table Part A Dalton's Law States That The Total Pressure, Ptotal, Of A Mixture Of Gases In A Container Equals The Sum Of The ...

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Question: A Review Constants 1 Periodic Table The Ideal Gas Law $PV = nRT$ V Part A Relates Pressure P. Volume V Temperature T And Number Of Moles Of A Gas, n. The Gas Constant Requals 0.08206 L.atm/(K.mol) Or 8.3145 J/(K.mol). The Equation Can Be Rearranged As Follows To Solve For: $PV = nRT$ This Equation Is Useful When Dealing With Gaseous Reactions Because Stoichiometric...

Solved: A Review Constants 1 Periodic Table The Ideal Gas ...

periodic law. the physical and chemical properties of the elements are periodic functions of their atomic numbers. periodic table. an arrangement of the elements in order of their atomic numbers so that elements with similar properties fall in the same column, or group.

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The Periodic Law . In 1869, Dmitri Mendeleev and Lothar Meyer individually came up with their own periodic law "when the elements are arranged in order of increasing atomic mass, certain sets of properties recur periodically." Meyer based his laws on the atomic volume (the atomic mass of an element divided by the density of its solid form), this property is called Molar volume.

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Short Answer. 117. In terms of the periodic law, explain which two of these elements are most similar: sodium (element 11), phosphorus (element 15), and sulfur (element 16). 118. What can you predict about the properties of xenon and helium, both in Group 18 in the periodic table? Why? Essay. 119.

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