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molecules produce a hydronium (H_3O^+) and a hydroxide (OH^-) ion $2 \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$. The ionization constant (K_w) of water is: $K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14} \text{ M}$ Acidic, Basic, & Neutral
If $[\text{H}_3\text{O}^+] > [\text{OH}^-]$ then solution is acidic.
If $[\text{H}_3\text{O}^+] < [\text{OH}^-]$ then solution is basic.

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