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Chapter 14 - Chemical Kinetics: Part 1 of 17

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Write the rate expression in terms of $\Delta [\text{reactant}] / \Delta t$ and $\Delta [\text{product}] / \Delta t$ for the reaction: (1) $\text{N}_2 + 3 \text{H}_2 \rightarrow 2 \text{NH}_3$. 2. Determine. The order of each reactant, The order of the reaction, Write the rate law and.

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Work in groups on these problems. You should try to answer the questions without referring to your textbook. If you get stuck, try asking another group for help. 1. For the reaction below, if substance A is disappearing at a rate of $(1.82 \times 10^{-2}) \text{ mol L}^{-1} \text{ s}^{-1}$, at what rate is C appearing? $[3 \text{A} + 3 \text{B} \rightarrow 5 \text{C} + 2 \text{D}]$ 2.

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Answers Kinetics Worksheet 1 - nathan.youshouldshare.me PDF Chemistry 12 Worksheet 1-1 - Measuring Reaction Rates Chemistry 12 Unit 1-Reaction Kinetics Worksheet 1-1 Measuring Reaction Rates Page 2 b) If the rate of consumption of magnesium is $5.0 \times 10^{-9} \text{ mol/s}$, find the rate of consumption of HCl in moles/s.

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Kinetics Worksheet With Answers

Exam 1 Worksheet Answers 3 Reaction order overall - 2 $k = 1.3 \times 10^{-9} \text{ M}^{-1} \text{ s}^{-1}$ a. $\ln [C]$ or $[D]$ versus time b. fast (relative to last example) c. $C + D \rightarrow E + B$ 5. 0.967 day (23.2 h) a. 8.38 g b. 0.717 day $^{-1}$ c. 19.3 days 6. 771 days (2.1 y) e. 0.234 g

Exam 1 Worksheet Answers

where k is the rate constant and c is the initial concentration. This is the equation that describes first-order kinetics. If $1/\text{concentration}$ vs. time is linear, the data will fit to the following equation: $1/[c] = kt + 1/[c]_0$. where k is the rate constant and c is the initial concentration.

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