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Revision notes & Important Questions of Thermodynamics

Solution: $\text{N}_2(\text{g}) + 3\text{H}_2 \rightarrow 2\text{NH}_3(\text{g})$ $\Delta_s = 2 - 4 = -2 < 0$. So entropy is negative. Hence option (c) is the answer. 3. If an endothermic reaction is non-spontaneous at freezing point of water and becomes feasible at its boiling point, then. (a) ΔH is -ve, ΔS is +ve. (b) ΔH and ΔS both are +ve.

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