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 points) $8 \text{ H}^+ (\text{aq}) + 4 \text{ Cl}^- (\text{aq}) + \text{MnO}_4^- (\text{aq}) \rightarrow 2 \text{ Cl}_2 (\text{g}) + \text{Mn}$

$3+ (\text{aq}) + 4 \text{ H}_2 \text{O} (\text{l}) \text{Cl}_2 (\text{g})$ can be generated in the laboratory
 by reacting potassium permanganate with an acidified solution of
 sodium chloride. The net-ionic equation for the reaction is given
 above. (a) A 25.00 mL sample of 0.250 M

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 (continued) (d) Calculate the value of the equilibrium constant, K_{eq} ,
 for the cell reaction at 25°C. Explain what the magnitude of K_{eq}
 tells you about the extent of the reaction. $50.72 \log 61 \text{ eq}$
 $0.0592 \text{ } 0.0592 \text{ nE } K \times == = K_{\text{eq}} = 6.5 \times 10^{60}$ Because the
 magnitude of K_{eq} is very large, the

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 (continued) (d) Calculate the value of the equilibrium constant,
 K_{eq} , for the cell reaction at 25°C. Explain what the magnitude of
 K_{eq} tells you about the extent of the reaction. $\log K_{\text{eq}} = . \text{ nE } 5 \times$
 $0.72 = = 61 \text{ } 0.0592 \text{ } 0.0592. K_{\text{eq}} = 6.5 \times 10^{60}$ Because the
 magnitude of K_{eq} is very large, the extent of the cell reaction is
 also very large and the reaction goes essentially to completion.

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