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III. THE SCHRÖDINGER EQUATION

Microscopic particles, like electrons, whose λ 's are on the order of their environment do not obey classical equations of motion. Electrons must be treated like waves to describe their behavior. 1927 Erwin Schrödinger wrote an equation of motion for particles (like electrons) that

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 $3 \text{ H}_2 + \text{N}_2 + 3 \text{ O}_2(\text{l}) \rightarrow 6 \text{ H}_2\text{O}(\text{l}) + 12 \text{ CO}_2(\text{g}) + 10 \text{ H}_2\text{O}(\text{g}) + \text{O}_2(\text{g})$ • 4 liquid molecules are converted into 29 gaseous molecules. • Increasing the number of moles increases the number of potential arrangements.

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Theory, Isotopes, Atomic Mass, Empirical & Molecular Formulas) Ch. 2 pg.35-48 · Lecture 2 Atomic Theory II (Waves, Light, Electronic Structure, Quantum Numbers) Ch. 6 pg. 187-218 · Lecture 3 Atomic Theory III (The Orbital, Rules of Filling Orbitals, Atomic & Ionic Radii ...

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In this video lecture, Chemguy finishes electron configurations and takes up periodic table trends. He covers essential periodic table trends, such as Ionization Energy, Electron Affinity, Electronegativity and Atomic Radius.

Lecture 7: Atomic Theory VII: Periodic Trends ...

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