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Chapter 10 Chemical Chemistry (12th

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Section 10.1 – The Mole: A Measurement of Matter. You often measure the amount of something by count, by mass, or by volume. A mole (mol) of a substance is 6.02×10^{23} representative particles of that substance. 6.02×10^{23} is called Avogadro's number. $1 \text{ mole} = 6.02 \times 10^{23}$ representative particles

Chapter 10 - Chemical Quantities

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- The basic unit that is used to determine the amount of a chemical substance is called a mole
- A mole (mol) of a substance is equivalent to 6.02×10^{23} particles of that substance
- The mole was founded by a scientist named Avagadro, and he decided to use the

CHAPTER 10: Chemical Quantities

Chapter 10 Chemical Quantities91. SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287–296) This section defines the mole and explains how the mole is used to measure matter. It also teaches you how to calculate the mass of a mole of any substance. Measuring Matter

(pages 287–289)

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