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According to modern atomic theory, we cannot be sure of the exact location of electrons in an atom. We predict that electrons will be relatively close

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N=3, L=1, M=1, M=1/2. This electron is found within the 2p orbital, which has the dumbbell shape, primarily based on the z axis. This electron would share this orbital with another electron, but would have a negative spin.

H-Atom Analogy Lab Discussion Questions | Atomic Orbital ...

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(d) If we count the number of dots in each spherical shell, we obtain the total probability of finding the electron at a given value of r. Because the surface area of each shell increases more rapidly with increasing r than the electron probability density decreases, a plot of electron probability versus r (the radial probability) shows a peak. This peak corresponds to the most probable radius for the electron, 52.9 pm, which is exactly the radius predicted by Bohr’s model of the hydrogen ...

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