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1 a) The relative atomic mass of an element is the weighted average mass of the isotopes of the element. (1) It is measured on a scale on which a carbon-12 atom has a mass of exactly 12 units. (1). b) The total mass of 100 atoms = $(75 \times 35) + (25 \times 37) = 3550$ (1) The average mass of 1 atom = $3550/100 = 35.5$ (1)

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Chemistry 4. a) The ammonia and hydrogen chloride particles have to diffuse through the air in the tube, colliding with air particles all the way. b) (i) Its particles will move faster. (ii) It would take slightly longer for the white ring to form, because the gas particles would be moving more slowly at the lower temperature.

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