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~~Mass | How to Pass Chemistry The Mole: Avogadro's Number and Stoichiometry Introduction to Combustion Analysis, Empirical Formula \u0026amp; Molecular Formula Problems Dalton's Law of Partial Pressure Problems \u0026amp; Examples - Chemistry Stoichiometry Mole to Mole Conversions - Molar Ratio Practice Problems Step by Step Stoichiometry Practice Problems | How to Pass Chemistry Balancing Chemical Equations Practice Problems Concept of Mole | Avogadro's Number | Atoms and Molecules | Don't Memorise~~

Chemistry B Moles Packet Answer

CHEMISTRY B- MOLES PACKET NAME: _____ HR: _____ PAGE 9 CHEMISTRY WORKSHEET # 6 MIXED MOLE PROBLEMS (GRAMS, MOLECULES, AND

LITERS) You now know three things a mole can be: a molar mass, 6.02×10^{23} molecules and, for a gas, 22.4 liters at STP.

CHEMISTRY B- MOLES PACKET

NAME: HR: _____ PAGE 1

CHEMISTRY B- MOLES PACKET NAME: _____ HR: _____ PAGE 4 CHEMISTRY WORKSHEET # 2: THE MOLE AS A UNIT OF MASS Now that you know how to find the mass of one mole of a substance (molar mass) you can easily find the mass of several moles or the mass of a fraction of a mole using the factor-label technique.

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A mole of carbon contains 6.02×10^{23}

atoms of carbon. A mole of helium contains 6.02×10^{23} atoms of helium. A mole of sodium contains 6.02×10^{23} atoms of sodium. A mole of gold contains 6.02×10^{23} atoms of gold. However, we have learned about seven elements that exist as diatomic molecules— H_2 , N_2 , O_2 , F_2 , Cl_2 , Br_2 , and I_2

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CHEMISTRY B- MOLES PACKET NAME: _____ HR: _____ PAGE 4
CHEMISTRY WORKSHEET # 2: THE MOLE AS A UNIT OF MASS
Now that you know how to find the mass of one mole of a substance (molar mass) you can easily find the mass of several moles or the mass of a fraction of a mole using the factor-label technique.
 $1 \text{ mol} = \text{a molar mass of an atom/molecule (g/mol)}$
EXAMPLE: What is the mass of 5.00 moles of water H_2O ?
STEP 1: $2 \text{ H} = 2 \times (1.01) = 2.02$
STEP 2: $? \text{ grams } \text{H}_2\text{O} = 5.00 \text{ moles } \text{H}_2\text{O} \times$

$18.02 \text{ g } \text{H}_2\text{O} = 90.10 \dots$

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CHEMISTRY B- MOLES PACKET NAME: _____ HR: _____ PAGE 6
CHEMISTRY WORKSHEET # 4 MOLAR MASSES AND AVOGADRO'S NUMBER
Now that you know two definitions of a mole (a molar mass and an Avogadro's number of particles) you can combine these two definitions into one problem.
EXAMPLE: How many molecules are there in 90.1 grams of water?
 $2 \text{ H} = 2 \times (1.01) = 2.02$
 $\text{O} = 1 \times (16.00) = 16.00$
 $\text{H}_2\text{O} = 18.02 \text{ g?}$
molecules $\text{H}_2\text{O} = 90.1 \text{ g } \text{H}_2\text{O} \times 1 \text{ mole } \text{H}_2\text{O} \times 6.02 \times 10^{23} \text{ molecules } \text{H}_2\text{O} = 3.01 \times 10^{24} \text{ molecules } \text{H}_2\text{O} \dots$

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Question: PAGES CHEMISTRY B-MOLES PACKET NAME: HR: CHEMISTRY WORKSHEET #5 MOLE PROBLEMS-MOLAR VOLUME OF A GAS We Have Learned That A Mole Is A Mass Of

Material And Number Of Particles. A Mole Can Also Be A Measure Of Volume When We Are Talking About Gases. You May Remember From Previous Science Classes That All Gases Behave Basically The Same As Far As The Physical ...

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