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Review Chemical Equilibrium Answers Section

CHAPTER 18 REVIEW Chemical Equilibrium CHAPTER 18 REVIEW Chemical Equilibrium SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. Write the equilibrium expression for the following hypothetical equation: $3A(aq) + B(aq) \rightleftharpoons 2C(aq) + 3D(aq)$ 2. a. Write the appropriate chemical equilibrium expression for each of the

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Figure 9.5.1: Equilibrium in reaction: $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$. Chemical equilibrium can be attained whether the reaction begins with all reactants and no products, all products and no reactants, or some of both. The figure below shows changes in concentration of H_2 , I_2 , and HI for two different reactions.

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Modern Chemistry 145 Chemical Equilibrium CHAPTER 18 REVIEW Chemical Equilibrium SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. ____ Raising the temperature of any equilibrium system always (a) favors the forward reaction. (b) favors the reverse reaction. (c) favors the exothermic reaction.

Review Chemical Equilibrium Section 2

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a state of balance in which the rates of the forward and reverse reactions are equal; no net change in the amount of reactants and products occurs in the chemical system (18.2) equilibrium position the relative concentrations of reactants and products of a reaction that has reached equilibrium; indicates whether the reactants or products are favored in the reversible reaction (18.2)

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Consider the following equation: $2\text{NO}(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g})$ At equilibrium, $[\text{NO}] = 0.80 \text{ M}$, $[\text{O}_2] = 0.50 \text{ M}$, and $[\text{NO}_2] = 0.60 \text{ M}$. Calculate the value of K . 4. Consider the following equilibrium equation: $\text{CH}_3\text{OH}(\text{g}) + 101 \text{ kJ} \rightleftharpoons \text{CO}(\text{g}) + 2\text{H}_2(\text{g})$. a. Increasing $[\text{CO}]$ will (a) increase $[\text{H}_2]$. decrease $[\text{H}_2]$. (c) not change [1-12].

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View Answer. The equilibrium system below shows the reaction of sulfur gas with methane to produce carbon disulfide and hydrogen gas. $\text{CH}_4(\text{g}) + 2\text{H}_2\text{S}(\text{g}) \rightleftharpoons \text{CS}_2(\text{g}) + 4\text{H}_2(\text{g})$; $\Delta H = +232.5 \text{ kJ}$...

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Solution: In systems 1 and 3, the concentration of A decreases from t_0 through t_2 but is the same at both t_2 and t_3 . Thus systems 1 and 3 are at equilibrium by t_3 . In system 2, the concentrations of A and B are still changing between t_2 and t_3 , so system 2 may not yet have reached equilibrium by t_3 .

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chemical equilibrium because its chemical equation is written with a double arrow like this. $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$ At equilibrium, the concentrations of reactants and products are constant, as you saw in Figures 18-2c and 18-2d. However, that doesn't mean that the amounts or concentrations of reactants and products are equal. That is seldom the case.

Chapter 18: Chemical Equilibrium

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