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5 The equilibrium between dichromate(VI) ions and chromate(VI) ions can be represented by  $\text{Cr}_2\text{O}_7^{2-}(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons 2\text{CrO}_4^{2-} + 2\text{H}^+$  (orange yellow) (a) At equilibrium, the expression for the equilibrium constant,  $K_c$ , is (1)  $A \quad K_c = \frac{[\text{CrO}_4^{2-}]^2}{[\text{Cr}_2\text{O}_7^{2-}][\text{H}_2\text{O}]}$  (2)  $B \quad K_c = \frac{2[\text{CrO}_4^{2-}]}{[\text{Cr}_2\text{O}_7^{2-}][\text{H}_2\text{O}]}$  (3)  $C \quad K_c = \frac{[\text{CrO}_4^{2-}]^2}{[\text{Cr}_2\text{O}_7^{2-}] + [\text{H}_2\text{O}]}$  (4)  $D \quad K_c = \frac{2[\text{CrO}_4^{2-}]}{[\text{Cr}_2\text{O}_7^{2-}] + 2[\text{H}_2\text{O}]}$

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