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If 3.0 L of a 5.2 M SrCl_2 solution is diluted to 50 L, what is the molarity of the diluted solution? To what volume should you dilute 30 mL of a 12.0 M H_2SO_4 solution to obtain a 0.170 M H_2SO_4 solution?

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Copyright © 2012 Nelson Education Ltd. Unit 1: Organic Chemistry U1-12 65. (a) The product of the

oxidation of cyclohexanol is cyclohexanone. (b) 66. (a) ketone: pentan-2-one (b) aldehyde: butanal (c) ketone: 4,5-dichlorohexan-3-one (d) aldehyde: 2,3-dimethylpentanal 67. Aldehydes and ketones both contain carbonyl groups. In an aldehyde, the carbonyl

Unit 1 Review, pages 120-127

Chemistry 2102C is the fourth course in the Chemistry Concentration in the Adult Basic Education program. Chemistry 1102, 2102A and 2102B are pre-requisites to Chemistry 2102C. The first part of this course introduces (or reviews) the basic concepts of solutions. It goes

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= 12.00 mmol - 5.40 mmol Unreacted ethanoic acid = 6.60 mmol Since 18.00 mL of NaOH(aq) was added to 20.00 mL of ethanoic acid solution, the total volume is now 38.00 mL. Solution: The concentration of unreacted ethanoic acid, $[\text{HC}_2\text{H}_3\text{O}_2(\text{aq})]$, can be determined: $23.2 \text{ g HC}_2\text{H}_3\text{O}_2(\text{aq}) \times \frac{1 \text{ mol}}{60.05 \text{ g}} = 0.386 \text{ mol}$ (aq) $23.2 \text{ g HC}_2\text{H}_3\text{O}_2(\text{aq}) \times \frac{1 \text{ mol}}{60.05 \text{ g}} = 0.386 \text{ mol}$ (aq) (6.60 mmol) = 38.00 mL