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Chapter 15 - Water and Aqueous Systems - 15.2 Homogeneous ...

The substance that dissolves in a solvent to form a solution. Solution. A homogeneous mixture. What substances dissolve most rapidly in water. Ionic compounds and polar covalent compounds. Solvation. The process by positive and negative ions of an ionic solid become surrounded by solvent.

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Section 15.2 Homogeneous Aqueous Systems 451 Solvated ions Surface of ionic solid molecules. The Solution Process Water molecules are in continuous motion because of their kinetic energy. When a crystal of sodium chloride is placed in water, the water molecules collide with it. Remember that a water molecule is polar, with a partial neg-

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A homogeneous equilibrium is one in which all of the reactants and products are present in a single solution (by definition, a homogeneous mixture). Reactions between solutes in liquid solutions belong to one type of homogeneous equilibria. The chemical species involved can be molecules, ions, or a mixture of both. For the homogenous reaction